

## SPECIFIC HEAT CAPACITY

$$q = c \times m \times \Delta T$$

**\*\*Use Appendix C Table C-6 to find specific heat capacities.\*\***

1. 5 g of copper was heated from 20°C to 80°C. How much energy was used to heat Cu?
2. If a 3.1 g ring is heated using 10.0 calories, its temperature rises by 17.9°C. Calculate the specific heat capacity of the ring.
3. If the temperature of 34.4 g of ethanol increases from 25°C to 78.8°C, how much heat has been absorbed by ethanol?
4. The temperature of a sample of iron with a mass of 10.0 g changed from 50.4°C to 25.0°C with the release of 114 J of heat. What is the specific heat of iron?
5. A 4.50 g nugget of gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25°C? The specific heat of gold is 0.129 J/g°C.